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| Science 9 | Unit B |
| Section 3.1 – 3.2: Naming Compounds |  |

Quiz Tomorrow

Chemistry Cards need to be handed in

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| Review Yesterday  NoteBook QUIZ  Dangers of DHMO  **Naming Compounds**  3 Versions   1. Common Name   What you see on a consumer label   1. Chemical Name   What you will see on a MSDS label   1. Chemical Formula   the elements are found in the chemical  (Same in all Languages)  Reading Chemical Formulas  Ni2O3(s)  2 Nickel Atoms 3 Oxygen Atoms Solid  Compounds and Elements Worksheet   1. **If A Metal is combined with a Non-Metal you get a Ionic Compound** 2. **If two Non-Metals are combined you have a Molecular Compound**   **Ionic Compounds**  Properties:   * Solids at room temperature * Good Conductors * Distinct Crystal Shape * Made from metals paired with non-metals   Pure Substances made from ‘ions’  Ions are ‘+’ or ‘-‘ charged elements  An element either gains or loses it’s electrons  **Naming Ionic Compounds**  Metals are named first, non-metals second  Metal Ions:   * ‘+’ charged * Check the periodic table for charges * If only 1 charge just the element name * If More than 1 charge then use roman numerals to show which charge is used   Non-Metal Ions:   * ‘-‘ charged * Check the periodic table * Non-metals only have one charge * Change the suffix to –ide   Polyatomic Ions   * ‘-‘ or ‘+’ Charged * Found on a periodic table * No need to change the name   Name to Formula  Calcium Chloride   |  |  | | --- | --- | | Metal  Ca2+  Balance Charges  2+  Ca | Non-Metal  Cl-   1. + 1-   Cl2 |   CaCl2(s)  Always solid unless in solution (water) | Table Salt, Water  Sodium Chloride, Dihydogen Monooxide  NaCl(s), H2O(l)  HF(g) = 1 Hydrogen 1 Fluorine Gas  Salt in Water = NaCl(aq)  C9H8O4(s) = 9Carbon 8Hydrogen 4 Oxygen Solid  Ionic = Metal + Non-Metal  NaCl melts at 801C  Forms Ions in water  Forms cubes  K+, Na+, Fe3+ or Fe2+  Cl-, Br-, SO4  Lithium Lead  Li+ Pb2+ or Pb4+  Lithium  Lead (II) or Lead (IV)  Oxygen Bromine  O2- Br1-  Oxide Bromide  Carbonate Sulfate  CO32- SO42-  Carbonate Sulfate |