

Chemistry 20	Science 10 Review
Lesson 7: Solubility of Ionic Compounds	84 mins

Solubility of Ionic Compounds

<ul style="list-style-type: none"> - Use the solubility table on pg. 6 of your data booklet. - Assume AQ for ambiguous ones... 	$\text{NaBr}_{(?)\text{(aq)}}$ $(\text{NH}_4)_2\text{CO}_{3(?)\text{(aq)}}$ $\text{AgCl}_{(?)\text{(s)}}$ $\text{Tc}(\text{HCO}_3)_{7(?)\text{(aq)}}$
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Single Replacement Reactions

<ul style="list-style-type: none"> - Compound reacts with a element to product a new compound and a new element. - Switch the like charges (eg Metals for metals) 	$\text{AB} + \text{C} \rightarrow \text{AC} + \text{B}$ <p>Examples</p> $(2)\text{NaCl}_{\text{(aq)}} + \text{Ca}_{\text{(s)}} \rightarrow (2)\text{Na}_{\text{(s)}} + \text{CaCl}_{2\text{(aq)}}$ $(\text{NH}_4)_2\text{Se}_{\text{(aq)}} + \text{F}_{2\text{(g)}} \rightarrow \text{Se}_{\text{(s)}} + (2)\text{NH}_4\text{F}_{\text{(aq)}}$
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Double Replacement

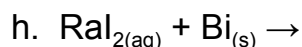
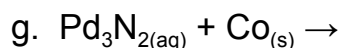
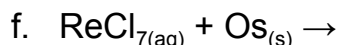
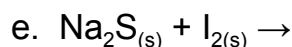
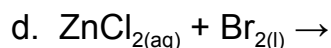
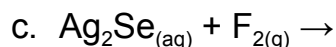
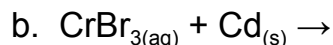
<ul style="list-style-type: none"> - Two compounds react to become new compounds 	$\text{AB} + \text{CD} \rightarrow \text{CB} + \text{AD}$ <p>Examples</p> $(2)\text{NH}_4\text{NO}_{3\text{(aq)}} + \text{BaCl}_{2\text{(aq)}} \rightarrow \text{Ba}(\text{NO}_3)_{2\text{(aq)}} + (2)\text{NH}_4\text{Cl}_{\text{(aq)}}$ $(3)\text{OsCl}_{4\text{(aq)}} + (2)\text{Sc}_2\text{O}_{3\text{(aq)}} \rightarrow (3)\text{OsO}_{2\text{(aq)}} + (4)\text{ScCl}_{3\text{(aq)}}$
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Chemistry 20 - Science 10 Review - Single and Double Replacement Reactions

Name: _____

Ensure that you include state subscripts. If the charge of a multivalent ion is ambiguous, use its *most common* form. Ambiguous states should be written as *aqueous*.

1. Complete and balance each of the following single replacement reactions.



2. Complete and balance each of the following double replacement reactions.

