

Chemistry 20	Unit 3
Lesson 6 - Acid Properties and Nomenclature	84 mins

Acids and Their Properties

<ul style="list-style-type: none"> - Corrosive - pH < 7 - Taste Sour - Electrolytes (produce ions in solution) 	$\text{HCl}_{(aq)}$ $\text{H}_3\text{PO}_{4(aq)}$ $\text{HNO}_{3(aq)}$ $\text{HNO}_{2(aq)}$
	ALL produce H^+ in water

Acid Nomenclature

<ul style="list-style-type: none"> - Names come from the suffix used when naming the "ionic" compound 	
<p>1) "-ide" becomes hydro_____ic acid</p> <p>H-ide a H-ick-y</p>	$\text{H}_3\text{N}_{(aq)}$ - Hydrogen nitride - hydron itric acid $\text{HAt}_{(aq)}$ - Hydrogen astatide - hydro astatic acid
<p>2) "-ate" becomes "_____ic acid" (no hydro)</p> <p>"ate" something "icky"</p>	$\text{H}_3\text{PO}_{4(aq)}$ - Hydrogen phosphate - phosphoric acid $\text{HNO}_{3(aq)}$ - Hydrogen nitrate - Nitric acid Oxalic Acid - Hydrogen oxalate - $\text{H}_2\text{OOC}(\text{COO})_{(aq)}$
<p>3) "-ite" becomes "_____ous acid"</p> <p>"B-ite" something "delici-ous"</p>	$\text{NHO}_{3(aq)}$ - Hydrogen nitrite - Nitrous acid Sulfurous acid - Hydrogen sulfite - $\text{H}_2\text{SO}_{3(aq)}$

Chemistry 20 - Unit 2 - Acid Nomenclature

Name: _____

1. Write formulas for each of the following acids.

a. hydrochloric acid.

b. hydrobromic acid.

c. hydroiodic acid.

d. sulfurous acid.

e. oxalic acid.

f. iodic acid.

g. perchloric acid.

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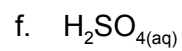
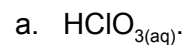
d. sulfurous acid.

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2. Write names for each of the following acids.



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